# Henderson-Hasselbalch equation

The Henderson and Hasselbalch equation describes the relationship between buffer composition and its acidity. Under simplifying assumptions:

$$\mathrm{pH} = \mathrm{p}K_A + \log \frac{c_\mathrm{B}}{c_\mathrm{A}},$$

where

pH is the resulting pH of the buffer,
pKA is the dissociation constant of the conjugate acid of the buffer,
$c_A$ and $c_B$ are the equilibrium concentrations of the conjugate acid and base forming the buffer.

## Links

### Related articles

- Derivation of the Henderson-Hasselbalch equation and other details.